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Tetrabutylammonium Iodide-Catalyzed Oxidative α-Azidation of β-Ketocarbonyl Compounds Using Sodium Azide

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Abstract

We herein report the oxidative α-azidation of carbonyl compounds by using NaN$_3$ in the presence of dibenzoylperoxide catalyzed by tetrabutylammonium iodide (TBAI). By utilizing these readily available bulk chemicals a variety of cyclic β-ketocarbonyl derivatives can be efficiently α-azidated under operationally simple conditions. Control experiments support a mechanistic scenario involving in situ formation of an ammonium hypoiodite species which first facilitates the α-iodination of the pronucleophile, followed by a phase-transfer catalyzed nucleophilic substitution by the azide. Furthermore, we also show that an analogous α-nitration by using NaNO$_2$ under otherwise identical conditions is possible as well.

Keywords

Introduction

Organic compounds containing an azide functionality are highly valuable synthesis targets that offer considerable potential for various applications and further manipulations [1-14]. For example, such molecules can be utilized to access free amines [3,13] and undergo Staudinger-type ligations [14]. Furthermore, they can be very efficiently employed for triazole-forming 1,3-dipolar cycloadditions with alkynes ("click-chemistry") [9-12]. As a consequence, the synthesis of organic azides is an important task and it comes as no surprise that a variety of conceptually complementary strategies to install azide groups in organic molecules have been reported [1-8]. α-Azido carbonyl derivatives are especially interesting targets which

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can be accessed by different approaches [6-8]. Maybe the most classical way to access organic azides is based on the utilization of pre-functionalized starting materials where a suited leaving group undergoes substitution using nucleophilic azide sources such as NaN_3 or $TMSN_3$ [6,7,15]. In addition, the recent years have seen remarkable progress in utilizing electrophilic azide-transfer reagents, i.e. hypervalent iodine-based compounds, for (asymmetric) α-azidations [16-23]. Besides these valuable approaches, which either require appropriate pre-functionalization of the starting materials (nucleophilic approach), or rely on more advanced N_3-transfer agents (electrophilic approach), over the course of the last years also α-azidations of enolate-type precursors using nucleophilic azide sources under oxidative conditions have been introduced very successfully [24-31]. Such oxidative coupling strategies of two inherently nucleophilic species allow for the direct utilization of simple starting materials in an efficient manner and especially the use of quaternary ammonium iodides as redox active catalysts has emerged as a powerful catalysis concept for such transformations [32-36]. These oxidative approaches, which usually proceed via the in situ formation of catalytically-competent ammonium hypoiodite species, can normally be carried out under operationally simple conditions, thus allowing for the use of easily accessible starting materials. Our group has a longstanding research interest in α-heterofunctionalization reactions under oxidative conditions [37-39] and we [30], as well as others [28,29,31], have recently also explored the use of simple quaternary ammonium iodides for oxidative α-azidations of carbonyl compounds (Scheme 1A). Hereby different strategies using different quaternary ammonium iodide derivatives and different azide sources were investigated and especially Ishihara's recent approach using NaN_3 in combination with the carefully designed achiral catalyst **C1** represents a remarkable advancement in this field (Scheme 1B [31]). In contrast to

previous oxidative quaternary ammonium iodide catalysis reports [28-30], this method does not require the use of TMSN$_3$, thus presenting an efficient oxidative α-azidation protocol utilizing NaN$_3$, which arguably represents the most easily available and cheapest nucleophilic N$_3$ source (for other remarkable approaches using alternative catalysts and oxidants see Ref. [24-26]). In addition to the racemic approach, they also showed that this reaction can be rendered enantioselective by using advanced Maruoka-type quat. ammonium iodides [40]. Interestingly, designer catalyst **C1** was found being catalytically superior compared to Bu$_4$NI (TBAI) when using H$_2$O$_2$ as the oxidant. Furthermore, it turned out that addition of PBN (phenyl N-t-butylnitrone) has a beneficial effect on the reaction and that carefully buffered conditions are best-suited. We have recently established the use of dibenzoylperoxide (DBPO) as a very powerful oxidant for oxidative heterofunctionalization reactions using simple nucleophilic inorganic salts as heteroatom transfer reagents [39,41]. This was successfully demonstrated for the non-catalyzed α-S(e)CN-functionalization of different pronucleophiles [39] as well as the benzylic azidation of alkylphenol derivatives with NaN$_3$ using TBAI as a catalyst [41]. Considering the fact that TBAI clearly represents one of the most easily available quat. ammonium iodides and keeping in mind our successfully demonstrated matching combination of this catalyst with NaN$_3$ and DBPO for our benzylic azidations [41], we were thus wondering if the use of these simple bulk chemicals also allows for the oxidative α-azidation of different carbonyl-based pronucleophiles. As outlined in this contribution, this reagent/catalyst system allows indeed for high yielding direct α-azidations of different (cyclic) β-ketocarbonyl derivatives (Scheme 1C), thus resulting in an operationally simple protocol to access α-azidated carbonyl derivatives. In addition, we have also carried out some test reactions using NaNO$_2$ instead of NaN$_3$ under otherwise identical conditions and

obtained a first proof-of-concept for the analogous, to the best of our knowledge so far unprecedented, quaternary ammonium hypoiodite-mediated α-nitration reaction.
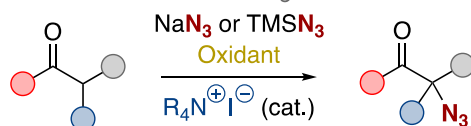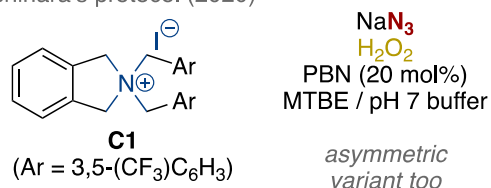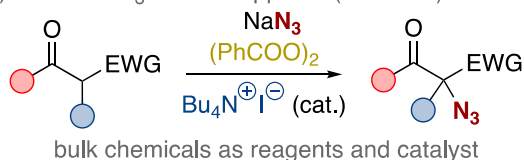

Scheme 1: General illustration of the oxidative α-azidation of carbonyl derivatives using quat. ammonium iodides (A), Ishihara's protocol using NaN$_3$ (B), and the herein reported combination of Bu$_4$NI (TBAI), NaN$_3$ and DBPO ((PhCOO)$_2$) (C).

Results and Discussion

We started our investigations by optimizing the α-azidation of the t-butyl-containing β-ketoester **1a** (Table 1 gives an overview of the most significant results obtained hereby). First experiments testing different oxidants in combination with Bu$_4$NI (30 mol%) in 1,2-dichloroethane (DCE), a solvent that we found to be well-suited for oxidative α-heterofunctionalizations before [39], showed that DBPO clearly outperforms all the other oxidants tested under these conditions (entries 1-5). While H$_2$O$_2$ gave **2a** in low yield only (entry 1), the use of mCPBA (entry 3) and t-BuOOH (entry 4) mainly resulted in the formation of the α-OH-ketoester **4**. On the other hand, Oxone performed significantly better (entry 2) but was also found to be inferior as

compared to DBPO, which allowed for the more or less quantitative "spot-to-spot" formation of **2a** without any noteworthy side-product formation (entry 5). Screening different catalyst / DBPO combinations next (entries 5-10), showed that the reaction requires a quat. ammonium halide containing an easily oxidizable counter anion, i.e. iodide or bromide (entries 5 and 7). No product formation was observed in the absence of any catalyst (entry 6) or in the presence of Bu_4NHSO_4 (entry 9) and the use of Bu_4NCl (entry 8) was not satisfying either. On the other hand, the beneficial effect of the quat. ammonium functionality was clearly underscored by employing KI instead of Bu_4NI (compare entries 10 and 5). While Bu_4NI allowed for the clean and selective formation of **1a**, we observed significant amounts of the α-I-ketoester **3** when using KI instead. Having established the combination of DBPO and Bu_4NI as the best-suited catalyst / oxidant combination for the α-azidation of **1a** using NaN_3, we finally optimized stoichiometry and catalyst loading (entries 11-15). Hereby we found the use of 1.2 eq. of NaN_3 with 1.2 eq. of DBPO and 20 mol% Bu_4NI as the best-suited and most economic reagent / catalyst combination, which allowed for the synthesis of **2a** in high isolated yield on 1 mmol scale as well (entry 14). Tests using other solvents under these optimized conditions were also carried out (details not given in the table), which showed that CH_2Cl_2 (95% NMR yield), toluene (95% NMR yield), acetonitrile (90% NMR yield), and THF (85% NMR yield) are also very well-tolerated.

Table 1: Optimization of the α-azidation of β-ketoester **1a**[a].
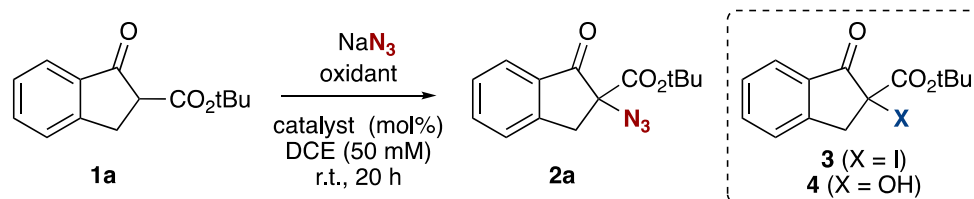
Entry	Cat. (mol%)	Eq. NaN$_3$	Oxidant (eq.)	**2a** (%)[b]	**3** (%)[c]	**4** (%)[c]
1	Bu$_4$NI (30)	2.2	H$_2$O$_2$ (2)	20	0	traces
2	Bu$_4$NI (30)	2.2	Oxone (2)	85	5	0
3	Bu$_4$NI (30)	2.2	mCPBA (2)	20	0	50
4	Bu$_4$NI (30)	2.2	tBuOOH (2)	30	traces	60
5	Bu$_4$NI (30)	2.2	DBPO (2)	>95	0	0
6	-	2.2	DBPO (2)	0	0	10
7	Bu$_4$NBr (30)	2.2	DBPO (2)	>95	0	0
8	Bu$_4$NCl (30)	2.2	DBPO (2)	15[d]	0	0
9	Bu$_4$NHSO$_4$ (30)	2.2	DBPO (2)	0[d]	0	0
10	KI (50)	2.2	DBPO (2)	60	30	0
11	Bu$_4$NI (30)	2.2	DBPO (1.2)	>95	0	0
12	Bu$_4$NI (30)	2.2	DBPO (0.5)	45	0	0
13	Bu$_4$NI (30)	1.2	DBPO (1.2)	>95	0	0
14	Bu$_4$NI (20)	1.2	DBPO (1.2)	95 (94)[e]	0	0
15	Bu$_4$NI (10)	1.2	DBPO (1.2)	85	0	0

[a] Unless otherwise stated, all reactions were carried out by stirring **1a** (0.1 mmol), the indicated amount of NaN$_3$, the catalyst, and the oxidant in 1,2-dichloroethane (DCE, 50 mM based on **1a**) at r.t. for 20 h.

[b] NMR yield using 1,3,5-trimethoxybenzene as an internal standard (given in 5% intervals).

[c] Determined by ^{1}H NMR of the crude product (given in 5% intervals).

[d] Complete conversion of **1a** but giving a rather complex reaction mixture.

[e] Isolated yield on 1 mmol scale.

Having identified high-yielding conditions for the synthesis of **2a**, we next carried out a series of control experiments in order to address the role of the catalyst's counter anion and the oxidant (Table 2). Running the reaction of **1a** and NaN_3 in the presence of stoichiometric amounts of Bu_4NI_3, I_2, Bu_4NIO_3, or Bu_4NIO_4 did not lead to any noteworthy levels of product formation (entries 1-4). In sharp contrast, the use of Bu_4NOH+I_2, which is known to give Bu_4NIO in situ [41-44], results in the formation of **2a** in a yield comparable to the above-described catalytic system. Accordingly, and in strong analogy to previous reports [41-43], the herein reported protocol most likely proceeds via in situ formation of a catalytically competent quat. ammonium hypoiodite species which then facilities the coupling of the two inherently nucleophilic reaction partners. To get further mechanistic insights we also carried out our standard reaction (Table 1, entry 14) in the presence of well-established radical traps like TEMPO, di-tert-butylhydroxytoluene (BHT), or 1,1-diphenylethene (DPE). In neither case any influence on the yield was observed, thus ruling out a mechanism involving radical species.

Table 2: Control experiments using different hypervalent iodine species[a].

Entry	Oxidant (1 eq.)	**2a** (%)[b]
1	Bu_4NI_3	0
2	I_2	15
3	Bu_4NIO_3	0

| 4 | Bu$_4$NIO$_4$ | 0 |
| 5 | I$_2$ / Bu$_4$NOH[c] | 95 |

Based on these mechanistic details obtained so far, we were also wondering if we could get any hints concerning possible reaction intermediates. Ishihara's group recently showed that their α-azidation protocol proceeds via in situ α-iodination first, followed by nucleophilic displacement by the azide [31]. Not surprisingly, when we analysed reactions shortly after the addition of all reagents we detected notable amounts of the α-iodinated β-ketoester **3** which then converted to the final product **2a** over time. Furthermore, we also synthesized compound **3** independently (by reacting **1a** with TBAI and additional KI in the presence of DBPO) and then resubmitted this compound to our ammonium salt-catalysed azidation reaction conditions, observing full conversion to **2a** as well. Considering all these details we thus propose a mechanistic scenario as outlined in Scheme 2. The catalyst gets oxidized to Bu$_4$NIO first, which then facilitates the α-iodination of **1a** (hereby either the formed benzoate or the hypoiodite itself may serve as a base). Intermediate **3** then undergoes a phase-transfer-catalysed nucleophilic substitution with NaN$_3$ thus delivering the final product **2a**.
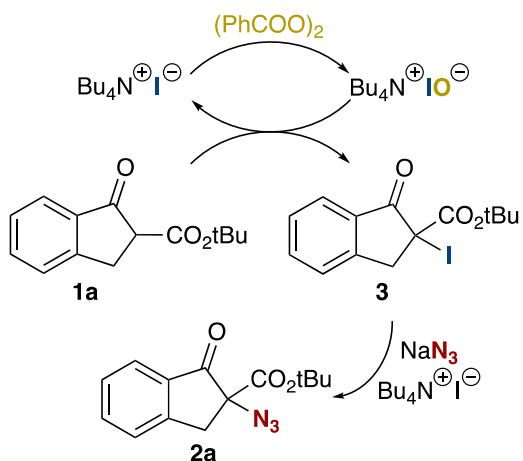
Scheme 2: Proposed mechanistic scenario.

With optimized conditions and a plausible mechanistic understanding at hand, we next investigated the application scope and limitations of this methodology. As outlined in Scheme 3, a series of differently substituted α-azide-β-ketoesters **2** as well as analogous α-N$_3$-β-ketoketones **5** and the α-N$_3$-β-ketoamide **6** could be accessed straightforwardly. Furthermore, this procedure was also successfully extended to γ-butyrolactone-based products **7**. Unfortunately, this methodology came to its limits when using tetralone-based β-ketoesters like compound **8**, which resulted in a complex product mixture, or the acylic β-ketoester **9**, which did not show any conversion under these conditions.
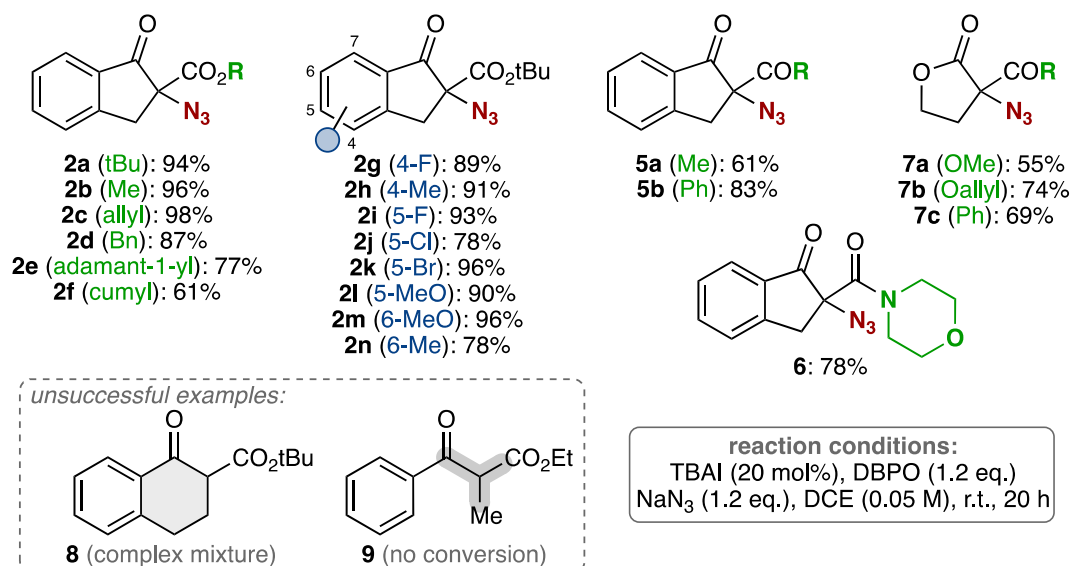
Scheme 3: Application scope.

Having established the TBAI / DBPO – mediated α-azidation using NaN3, we also briefly tested whether this concept can be extended to an analogous α-nitration approach. Different strategies for α-nitrations of carbonyl compounds have been reported [45-50], but the use of NaNO2 under simple oxidative conditions has so far received relatively little attention [51]. Gratifyingly, employing NaNO2 (2.2 eq.) under our established oxidative α-azidation conditions we found it possible to access the α-NO2-β-ketoesters **10a-c** as well (Scheme 4), which in our opinion represents an interesting proof-of-concept for an ammonium hypoiodite-mediated α-nitration. In this case we observed intermediate formation of the α-iodinated β-ketoester **3** as well (vide supra), which suggests an analogous mechanistic scenario as for the azidation (compare with Scheme 3). However, it should also be stated that the α-nitro products **10** were found to be rather instable compounds, which quickly underwent decarboxylation during column chromatography or upon prolonged reaction times. Also, tests with analogous β-ketoketones and β-ketoamides (compare with azidation products **5** and **6**, Scheme 4) did not give any products but resulted in the formation of

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a variety of unidentified side-products, thus illustrating that this α-nitration methodology seems to be less general than the α-azidation, which can most likely be attributed to the sensitivity of the products (containing a carbon with three EWGs).
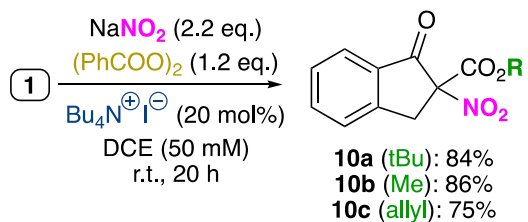
Scheme 4: Proof-of-concept for the analogous oxidative α-nitration.

Conclusion

α-Azidation reactions of carbonyl derivatives are powerful approaches to access valuable organic azides. In this contribution we report the direct α-azidation of cyclic β-keto-carbonyl compounds using NaN$_3$. This coupling of two inherently nucleophilic species is possible by carrying out the reaction under oxidative conditions using dibenzoylperoxide in the presence of a catalytic amount of tetrabutylammonium iodide (TBAI). Control experiments support a mechanistic scenario proceeding via in situ formation of a catalytically competent quaternary ammonium hypoiodite first. This higher oxidation state species then facilitates the α-iodination of the pronucleophile, followed by a phase-transfer catalysed nucleophilic substitution by the azide. Furthermore, we also obtained a first proof-of-concept for the conceptually analogous α-nitration by using NaNO$_2$ under otherwise identical conditions.

Experimental

General details

^{1}H-, ^{13}C- and ^{19}F-NMR spectra were recorded on a Bruker Avance III 300 MHz spectrometer with a broad band observe probe and a sample changer for 16 samples, on a Bruker Avance DRX 500 MHz spectrometer, and on a Bruker Avance III 700 MHz spectrometer with an Ascend magnet and TCI cryoprobe, which are all property of the Austro Czech NMR Research Center "RERI uasb". All NMR spectra were referenced on the solvent residual peak (CDCl$_3$: δ = 7.26 ppm for ^{1}H-NMR, δ = 77.16 ppm for ^{13}C-NMR, ^{19}F-NMR unreferenced). IR spectra were recorded on a Bruker Alpha II FTIR spectrometer with diamond ATR-module using OPUS software package. HRMS were recorded on an Agilent QTOF 6520 with an ESI source. Melting points are recorded using a Büchi M-560 apparatus and are reported uncorrected. TLC was performed on

Macherey-Nagel pre-coated TLC plates (silica gel, 60 F254, 0.20 mm, ALUGRAM®
Xtra SIL). Preparative column chromatography was carried out using
Davisil LC 60Å 70-200 MICRON silica gel. All chemicals were purchased from
commercial suppliers and used without further purification unless otherwise stated.

General α-Azidation Procedure

Sodium azide (7.8 mg, 120 µmol, 1.2 eq) and TBAI (7.4 mg, 20 µmol, 20 mol%) were
suspended in a stirred solution of the respective starting material (100 µmol, 1.00 eq)
in 1.0 mL of DCE at r.t.. Then, a solution of anhydrous dibenzoylperoxide (29.1 mg,
120 µmol, 1.2 eq) in 1.0 mL of DCE was added to the suspension and the mixture was
stirred for 20 h. The reaction solution was then diluted with 8 mL dichloromethane and
extracted with 5 mL of sat. aq. $NaHCO_3$. The aqueous phase was then extracted twice
with 10 mL of DCM. The organic layer and the extracts were then filtered consecutively
through a pad of anhydr. sodium sulfate and deactivated silica gel. The solvents were
removed in vacuo. In most cases the products were already obtained in sufficiently
high purity (> 95 %) after this work up. If necessary further purification by silica gel
column chromatography can be carried out.

Analytical details for the parent compound 2a: Obtained in 94% yield (25.7 mg,
94.0 µmol). *cf.*: 1.0 mmol scale, 94% yield (256.9 mg, 940.0 µmol; purified by
column chromatography on silica gel (eluent: heptanes/EtOAc = 19:1)). Yellowish-
white solid; Analytical data match those reported in literature [30]. ^{1}H-NMR (300 MHz,
$CDCl_3$, 298 K, $δ$ / ppm): 7.82 (d, $J = 7.7$ Hz, 1H), 7.66 (t, $J = 7.5$ Hz, 1H), 7.48 - 7.39
(m, 2H), 3.64 (d, $J = 17.2$ Hz, 1H), 2.99 (d, $J = 17.2$ Hz, 1H), 1.45 (s, 9H);
^{13}C-NMR (75 MHz, $CDCl_3$, 298 K, $δ$ / ppm): 198.1, 167.4, 152.3, 136.4, 133.3, 128.4,
126.5, 125.6, 84.6, 70.6, 38.6, 28.0; IR (neat, FT-ATR, 298 K, $\bar{v}$ / cm^{-1}): 2984, 2928,

2853, 2110, 1747, 1736, 1718, 1604, 1589, 1548, 1466, 1431, 1397, 1372, 1353, 1326, 1271, 1259, 1215, 1145, 1091, 1054, 1027, 961, 913, 871, 844, 834, 818, 804, 756, 729, 711, 688, 661, 623, 598, 561, 533, 459, 416; HRMS (ESI$^+$-QqTOF, m/z): calculated for $C_{14}H_{19}N_4O_3$ [M+NH$_4$]$^+$: 291.1452, found: 291.1452 (major); TLC (silica gel K60, 200 µm, F254, heptanes/EtOAc = 7:3, 298 K, R_f / 1): 0.64; MP: 65.0 - 67.5 °C.

General α-Nitration Procedure

Sodium nitrite (15.2 mg, 220 µmol, 2.2 eq) and TBAI (7.4 mg, 20 µmol, 20 mol%) were suspended in a stirred solution of the respective starting material (100 µmol, 1.00 eq) in 1.0 mL of DCE at r.t.. Then, a solution of anhydrous dibenzoylperoxide (29.1 mg, 120 µmol, 1.2 eq) in 1.0 mL of DCE was added to the suspension and the mixture was stirred for 20 h. The reaction solution was then diluted with 8 mL dichloromethane and extracted with 5 mL of sat. aq. NaHCO$_3$. The aqueous phase was then extracted twice with 10 mL of DCM. The organic layer and the extracts were then filtered consecutively through a pad of anhydr. sodium sulfate and deactivated silica gel. The solvents were removed in vacuo. In most cases the products were already obtained in sufficient purity (> 95 %) after this work up. If necessary further purification can be achieved by fast silica gel column chromatography (the products tend to decompose on silica gel).

Analytical details for the parent compound 10a: Obtained in 84% yield (23.3 mg, 94.0 µmol). white solid; ^{1}H-NMR (700 MHz, CDCl$_3$, 298 K, δ / ppm): 7.86 (d, $J = 7.7$ Hz, 1H), 7.71 (t, $J = 7.5$ Hz, 1H), 7.53 (d, $J = 7.7$ Hz, 1H), 7.48 (t, $J = 7.5$ Hz, 1H), 4.11 (d, $J = 17.9$ Hz, 1H), 3.99 (d, $J = 17.9$ Hz, 1H), 1.49 (s, 9H); ^{13}C-NMR (126 MHz, CDCl$_3$, 298 K, δ / ppm): 188.4, 162.0, 150.1, 137.0, 132.9, 129.1, 126.5, 126.2, 96.7, 86.1, 37.5, 27.8.; **IR** (neat, FT-ATR, 298 K, $\bar{v}$ / cm^{-1}): 2984, 2930,

2878, 2854, 1748, 1719, 1656, 1604, 1589, 1548, 1465, 1431, 1396, 1371, 1353, 1325, 1272, 1260, 1215, 1145, 1091, 1056, 1026, 961, 912, 871, 844, 834, 818, 803, 755, 730, 711, 688, 661, 625, 598, 561, 533, 459, 414; HRMS (ESI$^+$-QqTOF, m/z): calculated for $C_{14}H_{16}NO_5$ [M+H]$^+$: 278.1023, found: 278.1024; TLC (silica gel K60, 200 µm, F254, heptanes/EtOAc = 7:3, 298 K, R_f / 1): 0.47; MP: 75.9 – 78.4 °C.

Supporting Information

Full experimental and analytical details and copies of NMR spectra can be found in the online supporting information.

Acknowledgements

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